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or spontaneity of a given transformation. In this manner, chemical thermodynamics is typically used to predict the energy exchanges that occur in the following processes: Chemical reactions; Phase changes; The formation of solutions

Chemical thermodynamics - Wikipedia Given that the free energy of formation of liquid water is -237 kJ / mol, calculate the potential for the formation of hydrogen and oxygen from water. To solve this problem we must first calculate $\,$ G for the reaction, which is -2 (-237 kJ / mol) = 474 kJ / mol. Knowing that $\,$ G = -nFE o and n = 4, we calculate the potential is -1.23 V.

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PvT behavior of pure components; 3 — correlation and estimation of pure component properties; 4 — properties of mixtures; 5 — phase equilibria in fluid systems; 6 — caloric properties; 7 — electrolyte solutions; 8 — solid ...

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Wikipedia. Fundamentals of Thermodynamics.

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form a solid solution is a complicated matter
involving the chemical and crystallographic
properties of the substances as well as
temperature and pressure. At higher
temperatures substances have greater atomic
vibration and more open structures which
are easier to distort locally to accommodate
differently sized cations.

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