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School District
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ensures ...**

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Practice I 11.

How many grams
of potassium c
ride are needed
to make 200 ml,

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of a 2.5 M
solution?

(Answer: $2.5\text{M} =$

0.5 mol/L)

12. What is the

molarity of a

5.00×10^2 ml,

solution

containing 249 g

of calcium

iodide? (Answer:

1.69 M)

13. What is the

molarity of a

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**Solutions and
Molarity**

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Eventually, you

will enormously

discover a

further

experience and

talent by

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spending more
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always updates
the information
and also
understanding.

By doing this
can be only
exposed by
gathering the
new updates from
lots of sources.

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will utterly

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discover a new
experience and
achievement by
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cash. yet when?
get you agree to
that you require
to get those
every needs
bearing in mind
having
significantly
cash? Why don't
you try to

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acquire
something basic
in the
beginning?

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Molarity = 58.5

g = 1.69 M 0.562

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10. 78.9 g of $\text{CuSO}_4 \cdot 8\text{H}_2\text{O}$ is dissolved in 500.0 mL of water, calculate the molarity.

$$\begin{aligned} &78.9 \text{ g} \times 1 \text{ mole} \\ & \cdot \text{Molarity} = \\ &303.76 \text{ g} = 0.519 \\ &\text{M} \quad 0.5000 \text{ L} \cdot \end{aligned}$$

Stoichiometry

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W. J. Mouat

Chemistry 12

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it ends

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Hi Alondra, . . .
I'm sorry, but I
can't answer
your question,
because I do not
have the
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the molarity of

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the following
solutions given
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that: 1) 1.0
moles of
potassium
fluoride is
dissolved to
make 0.10 L of
solution. 2) 1.0
grams of
potassium
fluoride is
dissolved to
make 0.10 L of

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identifiers

_____ What does
molarity mean?

Number of moles
of solute. 1

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liter solution.
What is the
molarity of a
solution that
contains 4.53
moles of lithium
nitrate in 2.85
liters of
solution? 4.53
 $\text{mol LiNO}_3 =$
 $1.59 \text{ M LiNO}_3.$
 2.85 L soln

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Worksheet 2

**ANSWERS – Google
Docs**

Molarity = _____

Problems: Show
all work and
circle your
final answer. 1.
To make a 4.00 M
solution, how
many moles of
solute will be
needed if 12.0
liters of

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solution are
required? 2. How
many moles of
sucrose are
dissolved in 250
mL of solution
if the solution
concentration is
0.150 M? 3. What
is the molarity
of a solution of
HNO

Worksheet :

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Molarity

Molarity Name

Molarity+calcula
tions+(fillNinal
ltheboxes)+ ++so
lute+molesof+
solute+
grams+of+
solute+
volumeof++
solution+
Concentration+ (
Molarity,+M=mole
/L)+ ++NaCl+

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Molality

Osmolality

Osmolarity

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The answer is

1.00 mol/L.

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Notice that both the units of mol and L remain.

Neither cancels.

A symbol for mol/L is often used.

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KEY PART 1

Molarity: a
quantitative
description of
solution

concentration.

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of solute \div
liters of
solution

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all work and
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