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## Modern Chemistry Section Review Answers Chapter 28

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SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. Determine whether each of the following is an example of observation and data, a theory, a hypothesis, a

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control, or a model. observation and data a. A research team records the rainfall in inches

## 2 Measurements and Calculations

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## Study GuideChapter 5-21 Answer Key

CHAPTER 1 REVIEW Matter and Change SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. A horizontal row of elements in the periodic table is called a(n) period . 2. The symbol for the element in Period 2, Group 13, is B . 3. Elements that are good conductors of heat and electricity are metals . 4.

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1 Matter and Change - HUBBARD'S CHEMISTRY

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CHAPTER 5 REVIEW The Periodic Law SECTION 3 SHORT  
ANSWER Answer the following questions in the space provided. 1.  
c When an electron is added to a neutral atom, energy is (a) always  
absorbed. (c) either absorbed or released. (b) always released. (d)  
neither absorbed nor released. 2. d The energy required to remove  
an electron from a neutral atom is the atom ' s

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## 5 The Periodic Law

CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter

SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Explain the difference between the mass number and the atomic number of a nuclide. Mass number is the total number of protons and neutrons in the nucleus of an isotope.

## 3 Atoms: The Building Blocks of Matter

CHAPTER 6 REVIEW Chemical Bonding SECTION 1 SHORT

ANSWER Answer the following questions in the space provided. 1. a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis structures 2. b A covalent bond

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consists of (a) a shared electron.

## 6 Chemical Bonding

CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In a Stock system name such as iron(III) sulfate, the Roman numeral tells us (a) how many atoms of Fe are in one formula unit. (b) how many sulfate ions can be attached to the iron atom. (c) the charge on each Fe ion.

## 7 Chemical Formulas and Chemical Compounds

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Match description on the right to the correct crystal type on the left  
Modern chemistry chapter 10 review answers states of matter  
section 1. b ionic crystal (a) has mobile electrons in the crystal c  
covalent molecular crystal (b) is hard, brittle, and nonconducting a  
metallic crystal (c) typically has the lowest melting point of the four

Modern Chemistry Chapter 10 Review Answers States Of ...  
CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS  
Write the answer on the line to the left. Show all your work in the  
space provided. 1. 88% The actual yield of a reaction is 22 g and  
the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0  
mol of N<sub>2</sub> are mixed with 12.0 mol of H

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Use the following chemical equation to answer the question below:  
 $\text{Mg(s)} + 2\text{H}_2\text{O(aq)} \rightarrow \text{Mg}^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq}) + \text{H}_2(\text{g})$   
If 0.048 g of magnesium completely reacts in 20 s, what is the average reaction rate in moles/second over that time interval? Average rate 59.9 10 mol/s  
MODERN CHEMISTRY REACTION KINETICS 141

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