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Chemistry Chapter 4 Flashcards - Questions and Answers ...

See Section 4.1 (Page 103) . ANSWER: Correct Element Names, Symbols, and Atomic Numbers Learning Goal: To identify elements by their chemical symbol and their atomic number. Chemical elements are substances that are composed of only one kind of atom and cannot be broken down into simpler substances.

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NCERT Solutions for Class 11 Chemistry Chapter 4 Chemical ...

[Answer b\) is most likely as this molar mass corresponds to magnesium, which is a divalent metal, rather than to carbon or chlorine. Page 35 Ex 1.5.3 1. C 2. C 3. A 4. a\) 4.376x105, b\) 2.3x10-7, c\) 4.15x108, d\) 3.72x10-2, e\) 4.768x102, f\) 3.26x100 5. a\) 820000, b\) 0.00629, c\) 271 380 000 000, d\) 0.0000002, e\) 42, f\) 0.589](#)

ANSWERS - IB Documents

[Chemistry \(4th Edition\) Burdge, Julia Publisher McGraw-Hill Publishing Company ISBN 978-0-07802-152-7](#)

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[\(d\) 4. Answer. Answer: \(d\) 4 Explanation: Each water molecule can form a maximum of four hydrogen bonds with neighboring water molecules. The two hydrogens of the water molecule can form hydrogen bonds with other oxygens in ice, and the two lone pair of electrons on oxygen of the water molecule can attract other hydrogens in ice.](#)

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[Answer. Answer: \(c\) \(a\) and \(b\) both. Question 25. The activation energy in a chemical reaction is defined as. \(a\) the difference in energies of reactants and products. \(b\) the sum of energies of reactants and products. \(c\) the difference in energy of intermediate complex with the average energy of reactants and products.](#)

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[CHAPTER 4 REVIEW Arrangement of Electrons in Atoms SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. State the Pauli exclusion principle, and use it to explain why electrons in the same orbital must have opposite spin states. The Pauli exclusion principle states that no two electrons in an atom may have the](#)

4 Arrangement of Electrons in Atoms

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[Answer: According to VSEPR theory, if CH 4 were square planar, the bond angle would be 90°. For tetrahedral structure, the bond angle is 109°28'. For tetrahedral structure, the bond angle is 109°28'.](#)

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