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In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent:  $[\text{Au}(\text{CN})_2]^-$ ,  $\text{LaCl}_3$ , ethanol, and para-nitrophenol. When the limiting reactant is not apparent, we can determine which reactant is limiting by comparing the molar amounts of the reactants with their coefficients in the balanced chemical equation, just as we did in Section 11.4 .

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12.1 What is stoichiometry? In your textbook, read about stoichiometry and the balanced equation. For each statement below, write true or false. 1. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called

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