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donor. an acid, according to
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Hydrogen-ion acceptor. A base, according to the Bronsted-Lowry theory, ammonia acts as a base when it accepts hydrogen ions from water. Conjugate acid.

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turn blue litmus paper red,
taste sour, and react with
metals to produce hydrogen
gases. • Common acids
include vinegar ($\text{HC}_2\text{H}_3\text{O}_2$) and stomach acid (HCl). •
Bases turn red litmus paper
blue, have a bitter taste,
and are slippery to the
touch.

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Conjugate acid-base pair In

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the forward reaction, HCl releases a proton to form its conjugate base, Cl⁻. In the reverse direction, Cl⁻ accept a proton to form its conjugated base, HCl. A conjugated acid-base pair...

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HA(aq) + H₂O(l)

=equilibrium= H₃O⁺(aq) +

A⁻(aq) Acid 1 + Base 2 =

Acid 2 + Base 1. Explain

what happens in the

dissociation of an acid in

aqueous solution. The acid,

HA(aq), donates a proton to

water (which is therefore

behaving as a base), and

forms its conjugate base,

A⁻(aq).

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634 Chapter 18 • Acids and
Bases Section 118.18.1

Figure 18.1 Rhododendrons
flourish in rich, moist soil
that is moderately acidic,
whereas semper-vivum,
commonly called hen and
chicks, grows best in drier,
slightly basic soil.

Introduction to Acids and
Bases-!).)DEADifferent
models help describe the
behavior of acids and bases.

Chapter 18: Acids and Bases
Title: Chapter 20: Acids and
Bases 1 Chapter 20 Acids and
Bases 2 Acids. Give food a
sour flavor ; Aqueous

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solutions of acids are electrolytes ; React with compounds containing hydroxide ions to form salt and water; 3 Bases. Feel slippery ; Taste bitter ; Are electrolytes ; React with acids to form water and salt; 4 Formulas of Acids

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Chemistry Chapter 20 Acids and Bases. Acids. Bases. Acids. Base. Taste sour, turns litmus paper red, reacts with some metals to... Tastes bitter, feels slippery, turns litmus paper blue, reacts... compounds that form hydrogen ions when

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dissolved in water. A compound that produces hydroxide ions when dissolved in water.

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Bases 1 Chapter 10 Acids and
Bases. When we mix aqueous

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solutions of ionic salts, we are not mixing single components, but rather a mixture of the ions in the solid ; The ionic solid dissolves in the water ; We call a compound that dissolves in water soluble and if it doesnt, it is insoluble; 2 Electrolytes

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